**OXIDATION:** Loss of Electrons Atom 🡪 Ion

K(s) 🡪 K+  + e-

Mg(s) 🡪 Mg+2 + 2e-

**REDUCTION:** Gain of Electrons Ion 🡪 Atom

2Cl- + 2e- 🡪 Cl2(g)

O-2 + 2e- 🡪O2(g)

Lost electrons are written on the PRODUCT side (to the right of the arrow). OXIDATION

Gained electrons are written on the REACTANT side (to the left of the arrow). REDUCTION

OIL RIG—**O**xidation **I**s **L**oss **R**eduction **I**s **G**ain

Oxidation and reduction are PROCESSES.

An element will either be a REDUCING AGENT or an OXIDIZING AGENT.

STRONG Reducing agents: Groups 1 & 2 (3-13) {cations}

STRONG Oxidizing agents: Groups 16 & 17 (&15)